Gases and Vapor

1. "***STP***" means "***S***tandard ***T***emperature and ***P***ressure." (273 Kelvin & 1 atm)

2. ***Gases*** have widely-spaced particles that are in random motion.

3. ***Gases*** are easily compressed and have no definite shape or volume.

4. As the ***pressure*** on a gas increases, the ***volume*** decreases proportionally.

5. As the ***pressure*** on a gas increases, ***temperature*** increases.

6. As the ***temperature*** of a gas increases, ***volume*** increases.

7. *Always use Kelvins* for temperature when using the ***combined gas law***.

$$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$$

8. ***Real gas*** particles have volume and are attracted to one another, and thus do not always behave like ***ideal gases***.

9. Real gases behave more like ideal gases at *low pressures and high temperatures.*

USE THE REFERENCE TABLES!!!