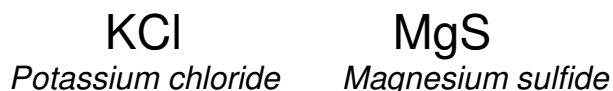


Chemical Formulas and Compounds

1. When naming binary ionic compounds, write the name of the positive ion (cation) first, followed by the name of the negative ion (anion) with the name ending in “-ide.” Example:



2. Polyatomic ions (Table E) are groups of atoms with an overall charge.
 NO_3^{1-} , NH_4^{1+} , SO_4^{2-} , etc.

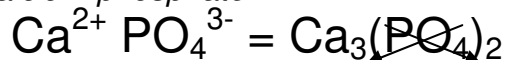
3. When naming compounds containing polyatomic ions, keep the name of the polyatomic ion the same as it is written in Table E.

Example:



4. Chemical formulas are written so that the charges of cations and anions neutralize one another.

Example: *calcium phosphate*:



5. The gram formula mass of a substance is the sum of the atomic masses of all of the atoms in it. **$\text{H}_2\text{SO}_4 = 98 \text{ g/mole}$**

$2 \times \text{H} = 2 \times 1 \text{ g/mole} = 2 \text{ g/mole}$ $1 \times \text{S} = 1 \times 32 \text{ g/mole} = 32 \text{ g/mole}$ $4 \times \text{O} = 4 \times 16 \text{ g/mole} = 64 \text{ g/mole}$	}	sum = 98 g/mole
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6. Know how to calculate the percentage composition of a compound. (Formula is on Table T.)

USE THE REFERENCE TABLES!!!