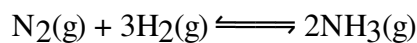


___ 9) What general type of reaction is illustrated in the diagram?

- 1) synthesis
- 2) decomposition
- 3) single replacement
- 4) double replacement

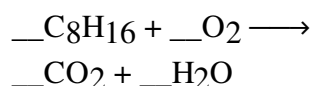
___ 10) Given the reaction:



What is the ratio of moles of $\text{H}_2(\text{g})$ consumed to moles of $\text{NH}_3(\text{g})$ produced?

- | | |
|--------|--------|
| 1) 6:6 | 3) 2:3 |
| 2) 3:2 | 4) 1:2 |

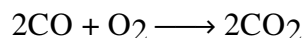
___ 11) When the equation



is correctly balanced using the *smallest* whole number coefficients, the coefficient of O_2 is

- | | |
|-------|-------|
| 1) 8 | 3) 1 |
| 2) 12 | 4) 16 |

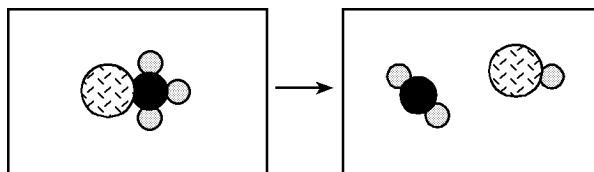
___ 12) Given the reaction:



What is the minimum number of grams of CO required to produce 88 grams of CO_2 ?

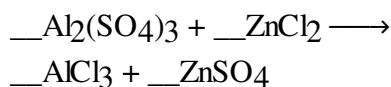
- | | |
|---------|---------|
| 1) 28 g | 3) 64 g |
| 2) 56 g | 4) 88 g |

___ 16) What general type of chemical reaction is illustrated in the particle diagram below?



- | | |
|-----------------------|-----------------------|
| 1) synthesis | 3) decomposition |
| 2) double replacement | 4) single replacement |

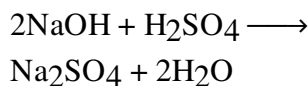
___ 13) When the equation



is correctly balanced using the *smallest* whole number coefficients, the sum of the coefficients is

- | | |
|------|------|
| 1) 5 | 3) 9 |
| 2) 8 | 4) 4 |

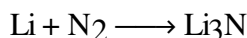
___ 14) Given the reaction:



What is the total number of moles of NaOH needed to react completely with 2 moles of H_2SO_4 ?

- | | |
|------|--------|
| 1) 1 | 3) 0.5 |
| 2) 2 | 4) 4 |

___ 15) Given the unbalanced equation:



When the equation is correctly balanced using *smallest* whole numbers, the coefficient of the lithium is

- | | |
|------|------|
| 1) 1 | 3) 3 |
| 2) 2 | 4) 6 |