

Acids and Bases

1. **Acids** and **bases** are both **good electrolytes**. Their solutions conduct electricity well.
2. Weak acids taste *sour*.
3. Weak bases taste *bitter*.
4. Acids and bases turn **indicators** different colors. They're listed on **Table M**.
5. Acids have a $\text{pH} < 7$.
6. Bases have a $\text{pH} > 7$.
7. **Tables K & L** list names and formulas of common acids and bases asked about on the Regents.
8. The metals above H_2 on **Table J** will react with acids to make H_2 gas bubbles.
9. **Arrhenius** says:
 - “Acids give off H^+ or H_3O^+ ions in solution.”
 - “Bases give off OH^- ions in solution.”
10. **Brønsted** says:
 - “Acids *donate* protons.”
 - “Bases *accept* protons.”
11. Acids and bases react in **neutralization** reactions to make **water** and a **salt**.
12. **Titrations** are controlled neutralization reactions used to find the concentration of an acid or base sample. Note the formula for it on Table T.

USE THE REFERENCE TABLES!!!